

Name of the Student: _____

Max. Marks : 15 Marks

Time : 15 Minutes

Mark Schemes

Q1.

C

The proportion of successful collisions increases because there is a decrease in the activation energy.

[1]

Q2.

C

The total number of molecules in the reaction mixture

[1]

Q3.

A

0.4

[1]

Q4.

C

The proportion of successful collisions increases because there is a decrease in activation energy.

[1]

Q5.

C

line C

[1]

Q6.

C

The proportion of successful collisions increases because the catalysed reaction has a lower activation energy.

[1]

Q7.

D

total number of particles present

[1]

Q8.
B

At a fixed temperature their average kinetic energy is constant

[1]

Q9.
C

They have no effect on the value of the equilibrium constant.

[1]

Q10.
D

There is an increase in the most probable energy of the molecules.

[1]

Q11.
B

[1]

Q12.
A

There are no molecules with zero energy.

[1]

Q13.
A

There is no reaction to form water because the molecules do not collide with sufficient energy.

[1]

Q14.
C

The rate of the reverse reaction increases.

[1]

Q15.
B

At a given temperature their average kinetic energy is constant

[1]