Practice Question Set For A-Level
Subject : Chemistry
Paper-2 Topic: Physical

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Name of the Student:_____
Max. Marks : 25 Marks **Time : 25 Minutes**

Mark Schemes

Q1.

Mark Range	The marking scheme for this part of the question includes an overall assessment for the Quality of Written Communication (QWC). There are no discrete marks for the assessment of QWC but the candidates' QWC in this answer will be one of the criteria used to assign a level and award the marks for this part of the question **Descriptor** an answer will be expected to meet most of the criteria in the level descriptor
4-5	– claims supported by an appropriate range of evidence – good use of information or ideas about chemistry, going beyond those given in the question – argument well structured with minimal repetition or irrelevant points – accurate and clear expression of ideas with only minor errors of grammar, punctuation and spelling
2-3	– claims partially supported by evidence – good use of information or ideas about chemistry given in the question but limited beyond this – the argument shows some attempt at structure – the ideas are expressed with reasonable clarity but with a few errors of grammar, punctuation and spelling
0-1	– valid points but not clearly linked to an argument structure – limited use of information or ideas about chemistry – unstructured – errors in spelling, punctuation and grammar or lack of fluency

(a) (i) M_r of $C_6H_5NH_2$ = 93 M_r of CH_3COCl = 78.5
total M_r of reagents = 264.5

1

$$\% \text{ atom economy} = \frac{M_r \text{ of wanted product}}{\text{total } M_r \text{ of all reagents}} \times 100 \text{ QWC}$$

1

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$= \frac{135}{\quad} \times 100 = 51.0\ \%$

1

(ii) expected yield = $\frac{10}{93} \times 0.5 \times 135 = 7.26$ kg

1

$\%\ \text{yield} = \frac{5.38}{7.26} \times 100 = 74.1\ \%$

1

(iii) Although yield appears satisfactory (74%) % atom economy is only 51% QWC

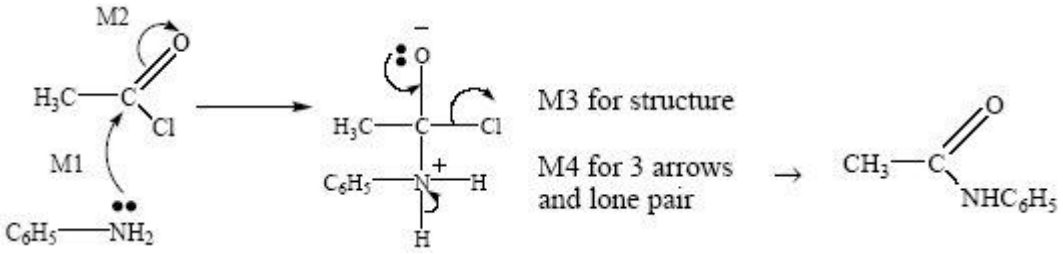
1

nearly half of the material produced is waste and must be disposed of QWC

1

(b) (nucleophilic) addition-elimination

1

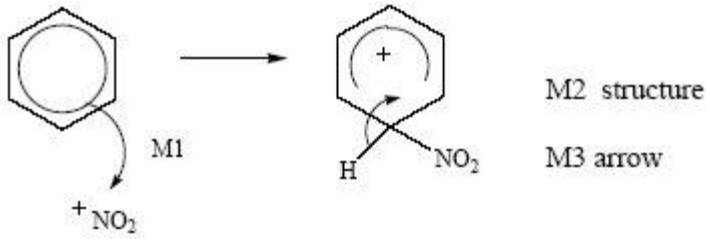


QWC (2)

4

(c) $HNO_3 + 2H_2SO_4 \rightarrow NO_2^+ + H_3O^+ + 2HSO_4^-$

1



3

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Q2.

(a) $Fe + H_2SO_4 \rightarrow FeSO_4 + H_2$

1

(b) $MnO_4^- + 8H^+ + 5Fe^{2+} \rightarrow Mn^{2+} + 4H_2O + 5Fe^{3+}$

1

(c) Moles MnO_4^- in 19.6 cm^3
= $19.6 \times 0.022 \times 10^{-3} = 4.312 \times 10^{-4}$

1

Moles Fe^{2+} in 25 cm^3
= $5 \times 4.312 \times 10^{-4} = 2.156 \times 10^{-3}$

1

Moles Fe^{2+} in 250 cm^3
= $10 \times 2.156 \times 10^{-3} = 2.156 \times 10^{-2}$

1

Mass Fe^{2+} = moles × A_r
$A_r = 2.156 \times 10^{-2} \times 55.8 = 1.203$ g

1

Percentage by mass of carbon
= (1.270 – 1.203) × 100/1.270
= 5.28%

1

(d) Repeat the titration and take an average of the concordant results

1

(e) Analyse several samples from different parts of the molten iron

1

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